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Ions are atoms with a charge. When ions of opposite charges are attracted, they form ionic bonds. Instead of electrons being shared, it is the electrostatic forces of attraction that hold ions together in ionic compounds, much like magnets sticking to one another.

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An ionic bond essentially donates an electron to the other atom participating in the bond, while electrons in a covalent bond are shared equally between the atoms. The only pure covalent bonds occur between identical atoms. Usually, there is some polarity (polar covalent bond) in which the electrons are shared, but spend more time with one atom than the other. Ionic bonds form between a metal and a nonmetal. Covalent bonds form between two nonmetals.

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